

CORRECTIONS TO
MOLECULAR THERMODYNAMICS OF FLUID-PHASE EQUILIBRIA

3rd EDITION - third printing
(PRENTICE-HALL, 1999)

Page	Location	Where is	Should be
xiv	3 lines up	<i>cognescere</i>	<i>cognoscere</i>
5	9 lines up	bot	not
16	Eq. (2-21)	Vdp	VdP
26	2nd reference	<i>Termostatics</i>	<i>Thermostatistics</i>
107	4 lines up	ψ	Ψ
118	Line 1	is 3×10^{-30} C m	is 1.08 debye
118	Line 2	2.63×10^{-30} m ³	2.60×10^{-30} m ³
160	Eq. (5-52)	0.866	0.886
174	Eq. (5-95)	$K = \frac{P_{A_2}}{P_A^2} = \frac{y_{A_2} P^0}{y_A^2 P}$	$K = \frac{(P_{A_2} / P^0)}{(P_A / P^0)^2} = \frac{y_{A_2} P^0}{y_A^2 P}$
176	Fig. 5-27	K, bar ⁻¹	K
205	2 reference up	<i>R. Soc.</i>	<i>Proc. R. Soc.</i>
207	Line 4	1.5	-1.5
209	Line 10	Add: $B_{22} = -12; B_{23} = -53$	
211	Line 7	Fluoride	fluoride
262	Label to Fig. 6-18	$(g_{ij} - g_{ii})$	$(g_{ij} - g_{jj})$
269	4 lines up	line (1)	line T_1
270	7 lines up	x'_1 and x''_1	x_1 and x_2
281	Eq. (6-154)	$(A'_{12} + A'_{13} - A'_{32}) \left(\frac{A'_{13}}{A'_{31}} \right)$	$\left(A'_{12} + A'_{13} - A'_{32} \frac{A'_{13}}{A'_{31}} \right)$
304	Table in Probl. 16	pressure bar)	pressure (bar)
304	Probl. 17	53.3 Estimate	53.3. Estimate
304	Probl. 18	4.7 At	4.7. At
310	Eq. (7-6)	$-\frac{a_1 x_1}{v_1^L}$	$= \frac{a_1 x_1}{v_1^L}$
310	Eq. (7-7)	$-\frac{a_2 x_2}{v_2^L}$	$= \frac{a_2 x_2}{v_2^L}$
370	Line 8	When $K = 0$, γ_1 for	When $K = 0$, $\gamma_1 = 1$ for
387	7 lines up	given by Eq. (7-216)	given by Eq. (7-221)
388	Eq. (7-244)	$\frac{E_o}{2kT} = \sum_{n=1}^4 \sum_{m=1}^M \frac{mA_{nm}}{\tilde{T}^n \tilde{v}^m}$	$\frac{E_o}{2kT} = \sum_{n=1}^4 \sum_{m=1}^M \frac{cA_{nm}}{\tilde{T}^n \tilde{v}^m}$
397	Eq. (7-246)	N_{Av}	N_A
409	6th reference	1998, <i>Fluid Phase Equilibria</i> , in press.	1999, <i>Fluid Phase Equilibria</i> , 162: 289.
412	Table in Probl. 3	(J cm ⁻³)	(J cm ⁻³) ^{1/2}
458	Fig. 8-19	• 1 bar ○ 1013 bar	○ 1 bar • 1013 bar
465	Eq. (8-95)	\tilde{v}	\tilde{v}
501	9 lines up	11: 192 (1978).	11: 192.
503	Line 6	33: 1047.	49: 2765.
503	10 lines up	When Φ is	When Φ_1 is
514	8 lines up	$\gamma_{\pm}^{(m)} = [(\gamma_{Ca^{2+}})(\gamma_{Cl^-})]^{1/3}$	$\gamma_{\pm}^{(m)} = [(\gamma_{Ca^{2+}})(\gamma_{Cl^-})^2]^{1/3}$

571	Caption to Fig. 9-28	■ : chymotrypsin	□ : chymotrypsin
593	5 lines up	changes in Y	changes in W
670	Table in Probl. 10	For n -hexane, interchange data for δ and ν	
708	Figure 12-24	R=OH	R=OCH ₃
		R=CH ₃	R=OH
731	5 lines after Eq.(12-81)	$K_B = K_{AB} = 0.$	$K_{B_2} = K_{AB} = 0.$
746	Table in Probl. 8	2 132	2.132
746	16 lines up	atm (cm ³ mol ⁻¹) K ^{1/2}	atm (cm ³ mol ⁻¹) ² K ^{1/2}
746	13 lines up	$a_{ij} = [a_i^{(0)} a_i^{(1)}]^{1/2}$	$a_{ij} = [a_i^{(0)} a_j^{(0)}]^{1/2}$
746	9 lines up	X	X
818	Eq. (G-2)	$n_B = \sum_i i n_{B_i} + \sum_i \sum_j i n_{A_i B_j}$	$n_B = \sum_j j n_{B_j} + \sum_i \sum_j j n_{A_i B_j}$
847	Line 5	data 112, 844	data 112, 845
854	Line 2	326	626
859	14 lines up	445, 705	444